LAB REPORT:
COMBINATION OF DETERMINATION OF MELTING POINTS AND
RECRYSTALLIZATION

A. Melting Point of Pure Compounds

1. Compound assigned: _____________________
2. Description of compound:

3. Literature melting point: ______
4. Melting point of above compound: ______

B. Melting Point of Impure Compounds

5. Identify of second compound: ________________
6. Melting range of mixture: __________

C. Recrystallization

7. Unknown assigned: ________________
8. Description of crude unknown (before recrystallization):

9. Mass of crude unknown obtained: _______
10. Mass of purified unknown obtained: _______

Name: _______________________________  Date: _____________
11. Description of purified unknown (after recrystallization):

12. Melting point of purified compound: __________

13. % of material recovered: _____________________


15. Melting point of unknown mixed with reference sample: ______

16. Literature melting point: ______

17. Explain how you determined the identification of your unknown. What proof do you have that your identification is correct?

Questions:

1.) Why do you use hot ethanol to rinse your flask during the first filtration but cold 1:1 ethanol-water during the second?

2.) Why is your hot solution cooled gradually and not immediately immersed in an ice bath?